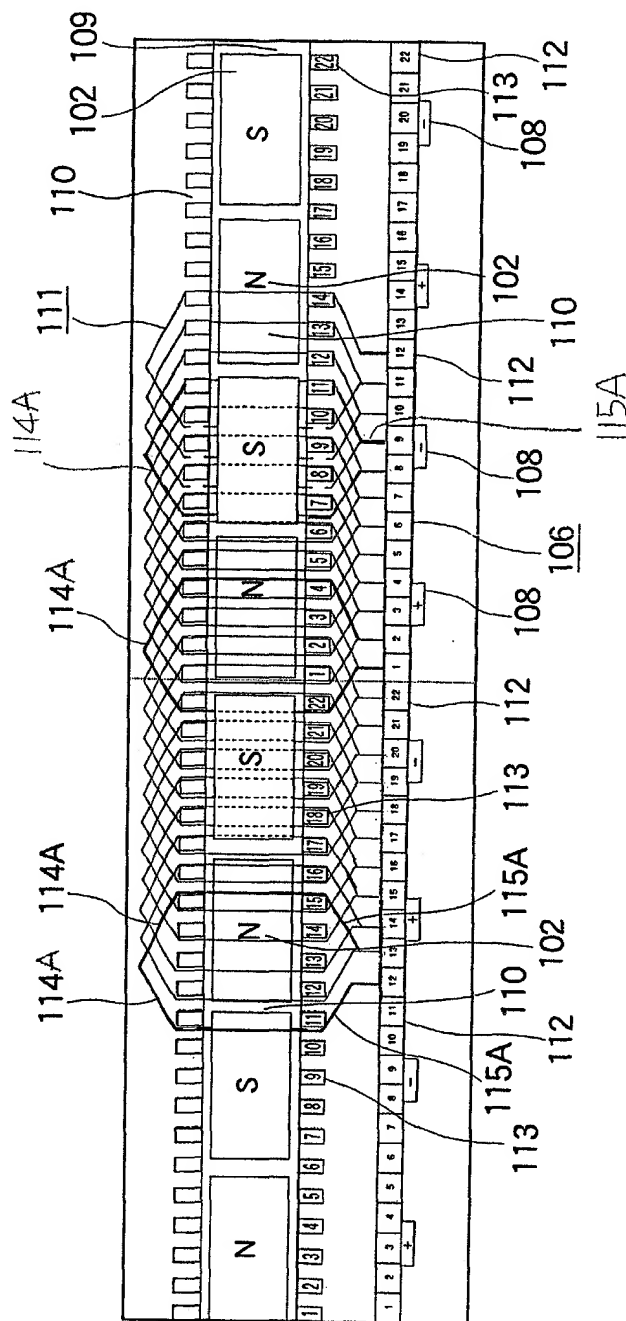


FIG. 2



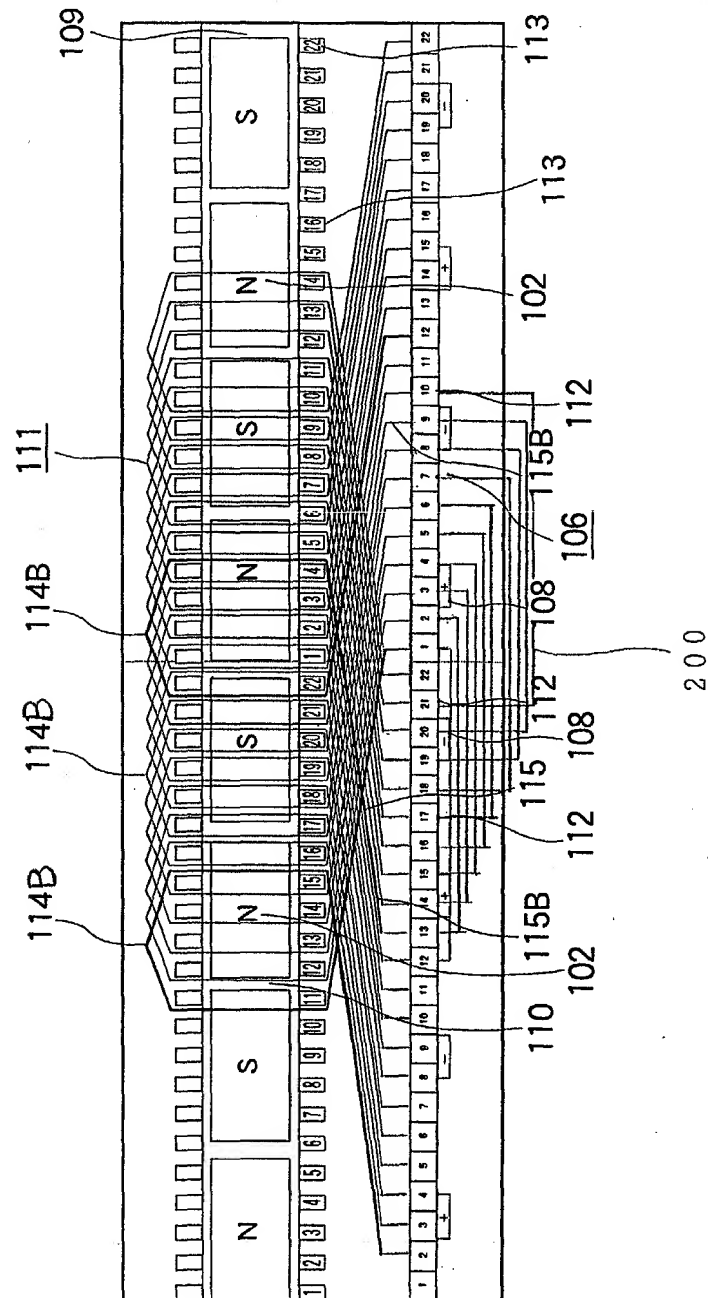
[illegible]

FIG. 4

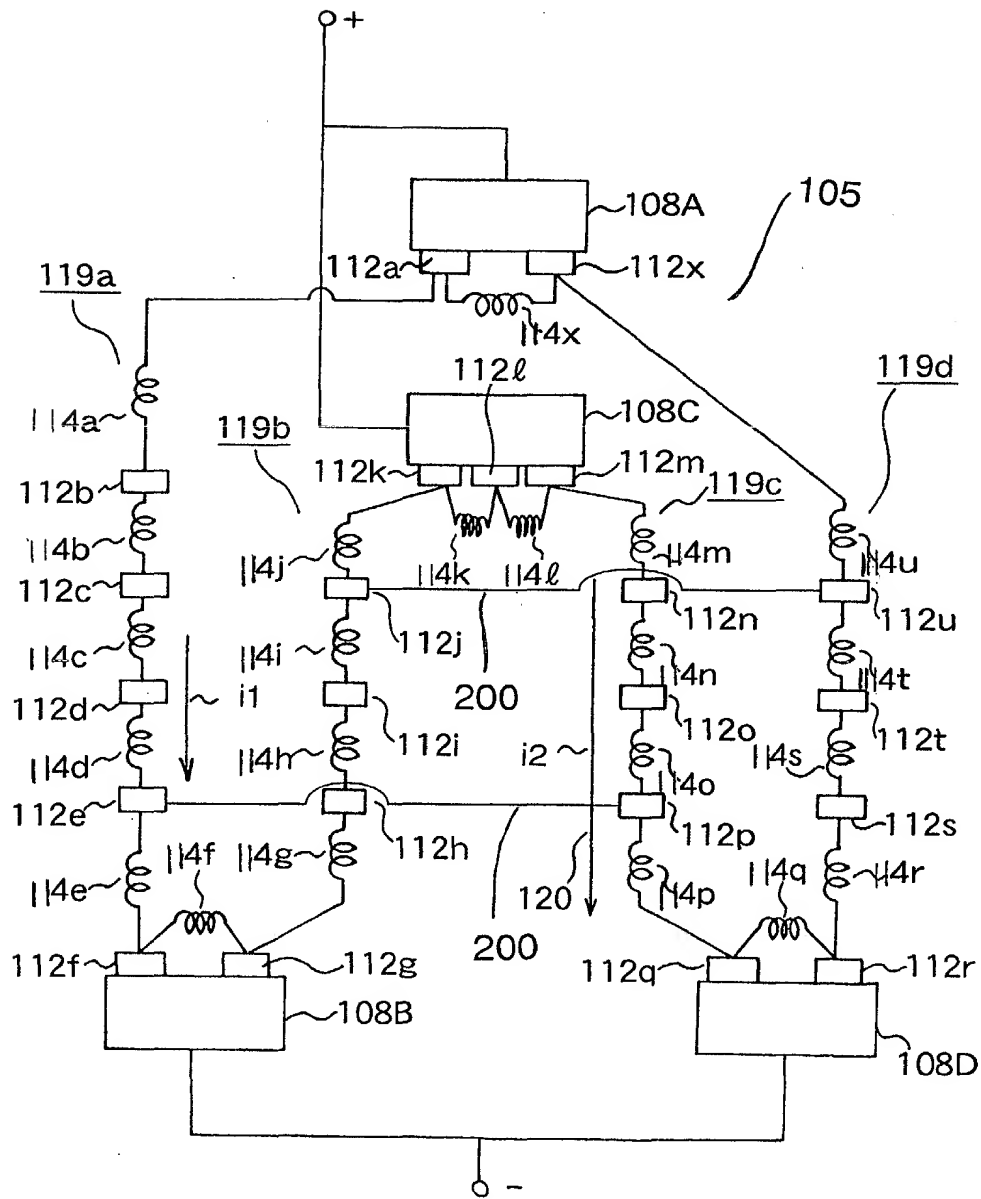
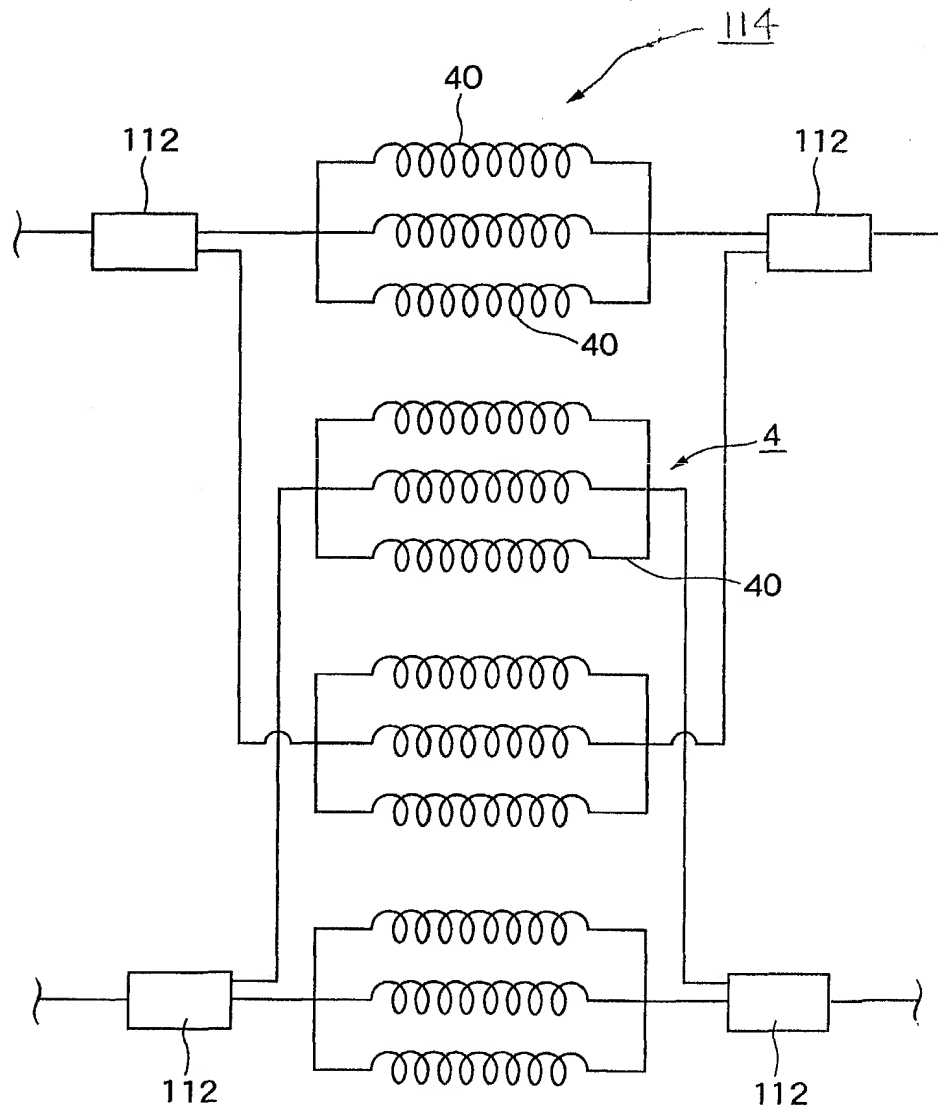


FIG. 5



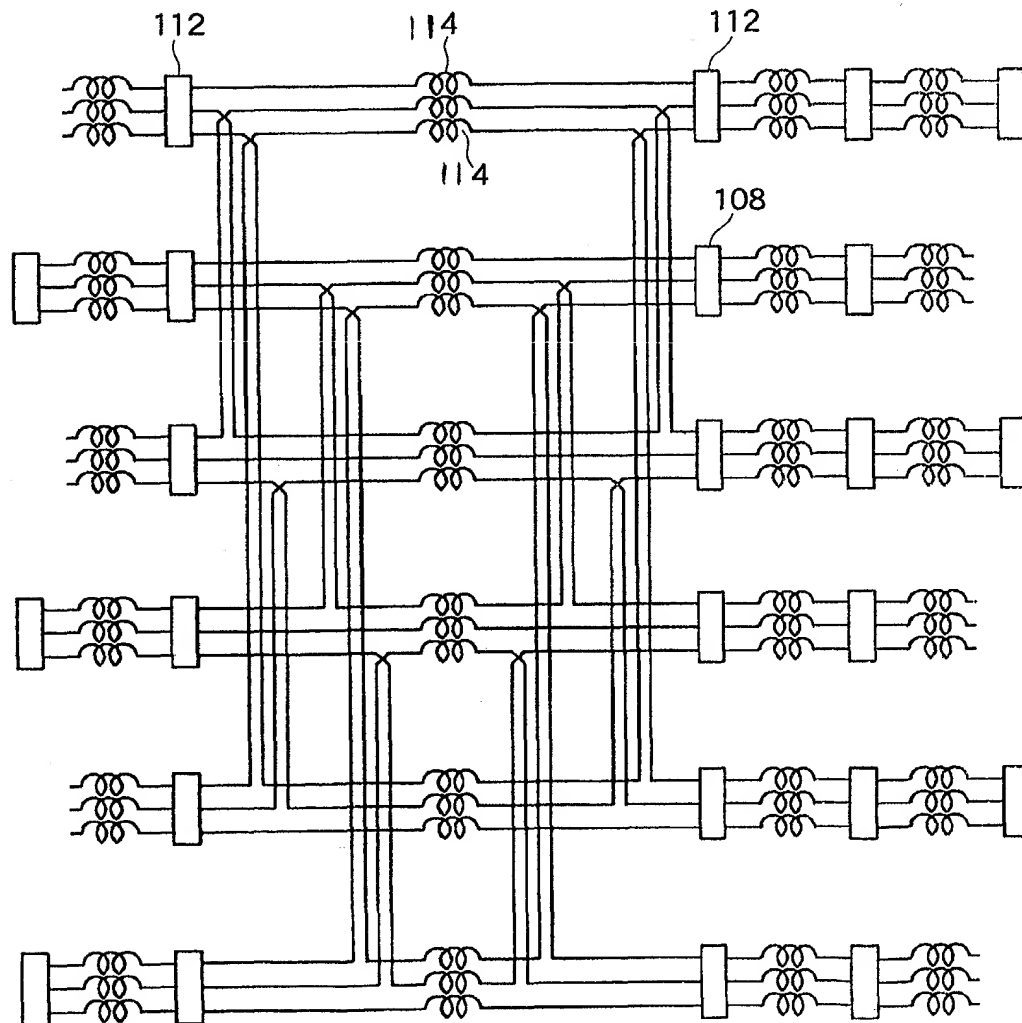
[illegible]

Figure 1

Diagram illustrating the experimental setup for measuring the effect of temperature on the rate of reaction between hydrogen peroxide and potassium iodide.

The diagram shows a test tube containing a mixture of hydrogen peroxide (H_2O_2) and potassium iodide (KI). The test tube is placed in a water bath maintained at a constant temperature ($T^\circ\text{C}$). The reaction proceeds, producing oxygen gas (O_2) and water (H_2O). The volume of oxygen gas collected over time is measured using a gas syringe or a graduated cylinder inverted in a trough of water.

The reaction equation shown is:

$$\text{H}_2\text{O}_2 + \text{KI} \rightarrow \text{I}_2 + \text{H}_2\text{O} + \text{O}_2$$

The diagram also indicates the measurement of the initial concentration of hydrogen peroxide (H_2O_2) and the initial concentration of potassium iodide (KI).

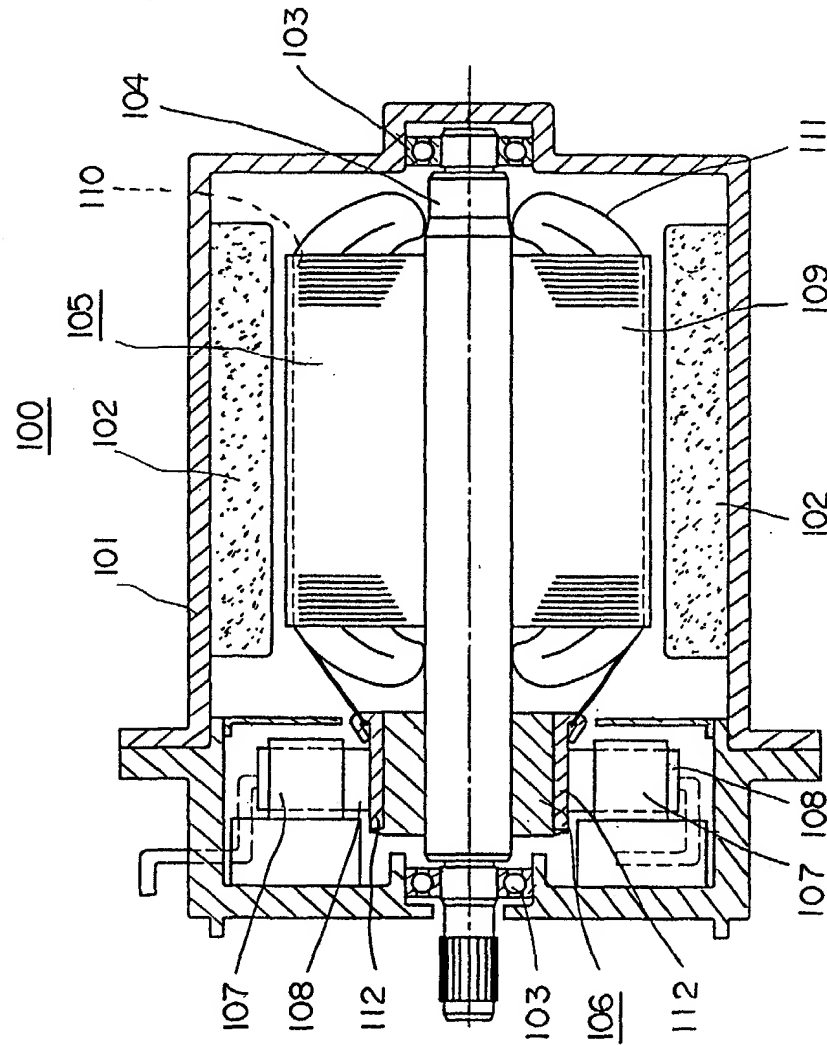


FIG. 8

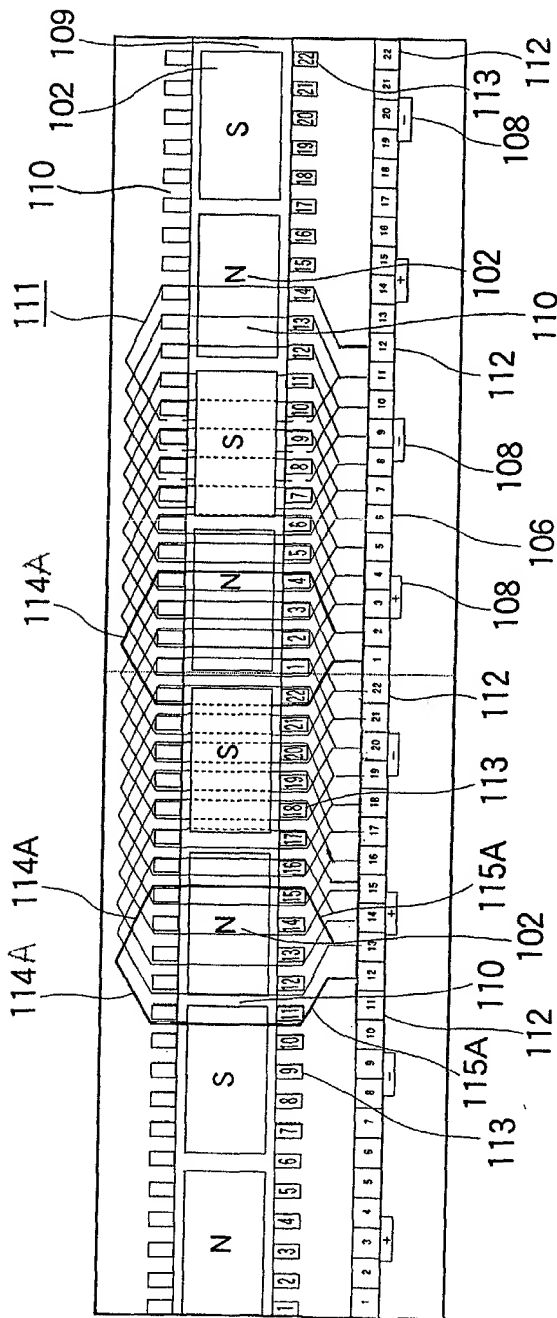


FIG. 9

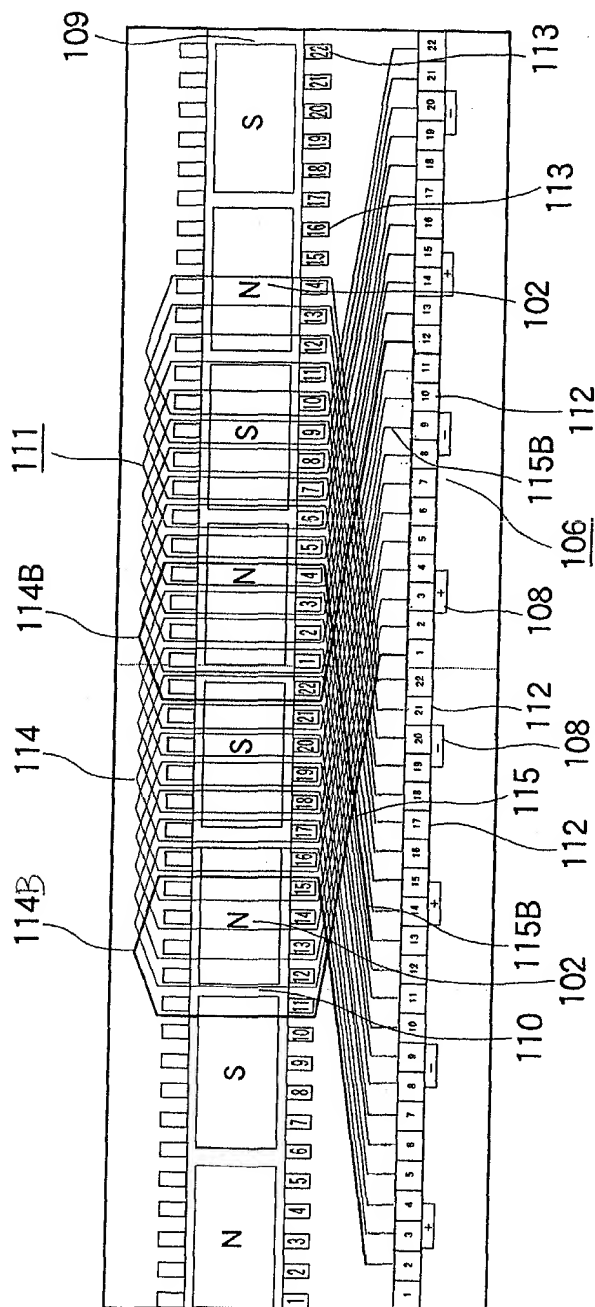


FIG. 10

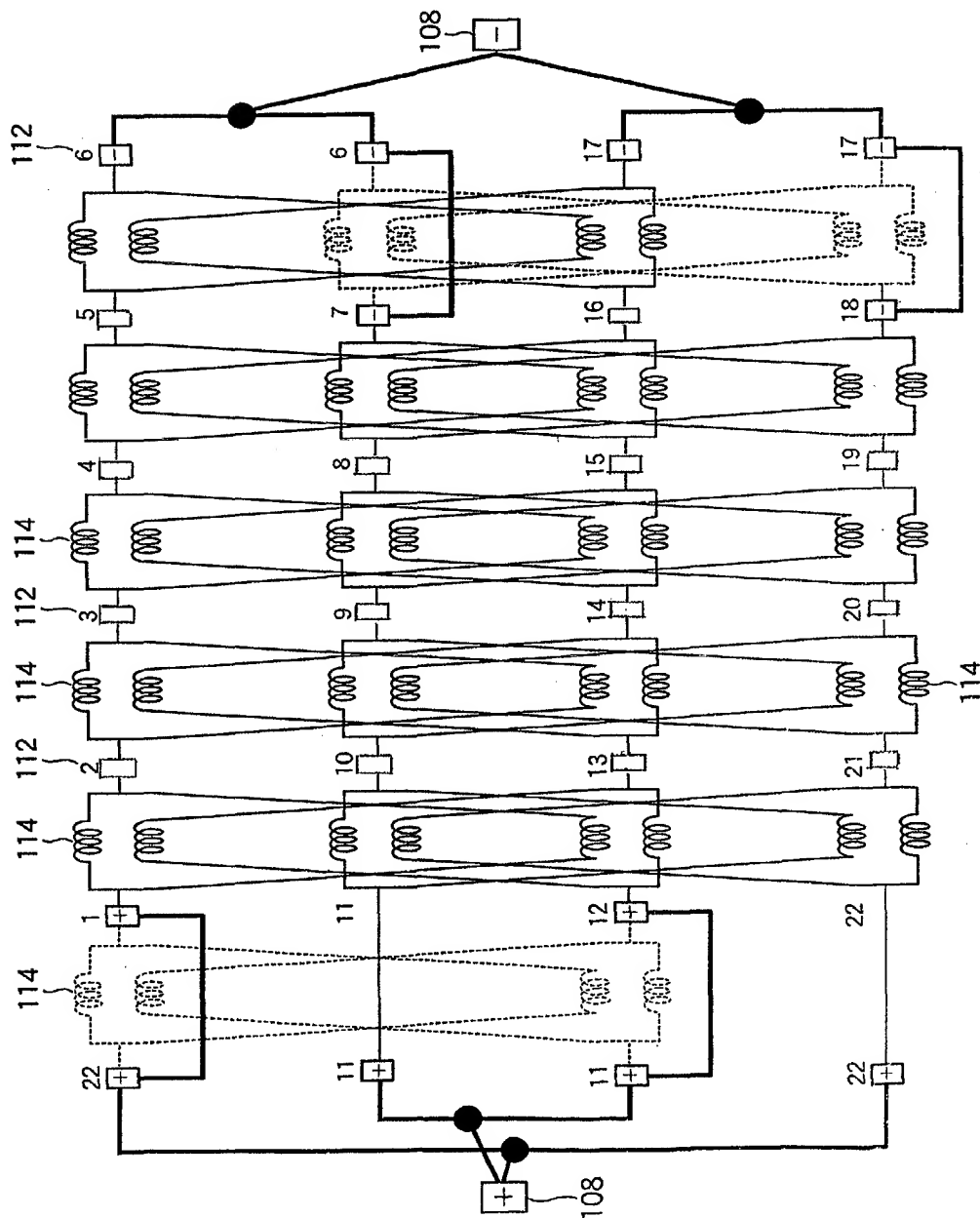


FIG. 11

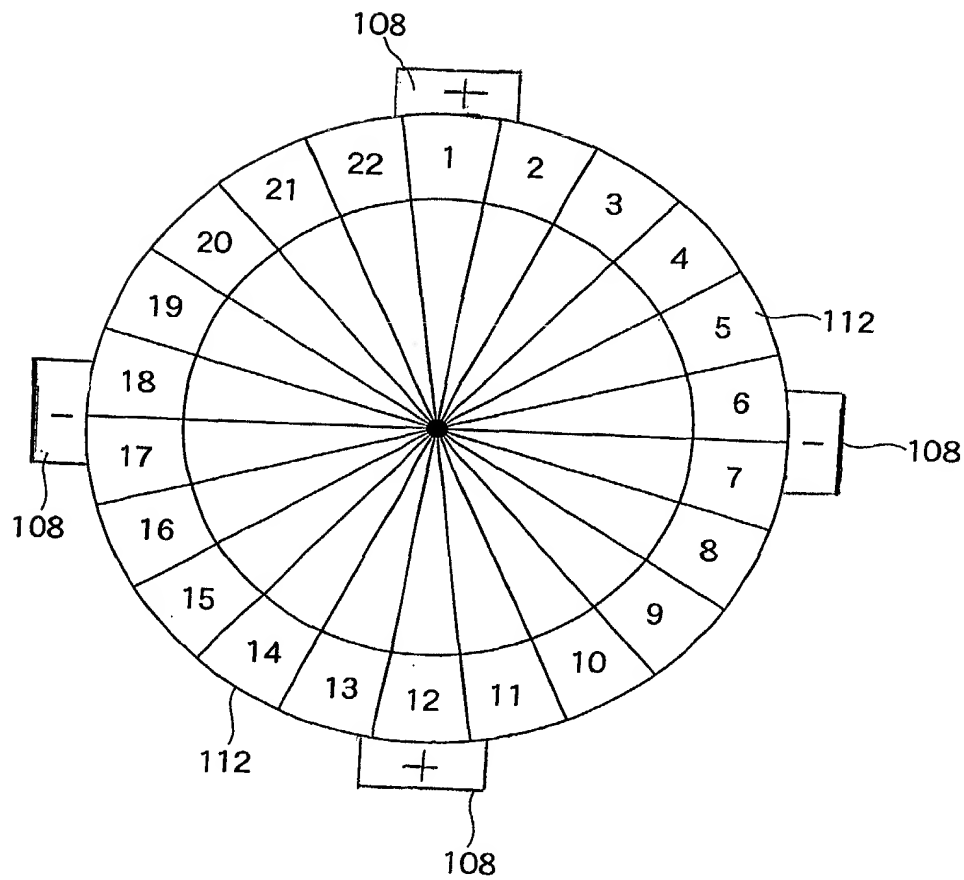


FIG. 12

